The $a_{1g}$ and $a_{2u}$ orbitals are much lower in energy than the $b_{1u}$ atomic orbitals, so the bonding orbitals are primarily centered on the rings.
2.5 O more EN than N
  $\sigma$ bonding - $a_1$ & $b_2$
  $\pi$ bonding - $b_1$

$a_2$: O group orbital - nothing of same symmetry as N

$\sigma\tau$ overall

$\pi$ bonding

LUMO

HOMO

primarily bonding from $p$ orbitals

primarily non bonding

$\sigma\tau$ bonding

from $s$ orbitals

N atomic orbitals

O group orbitals

1. Assume O group orbitals from $s$ orbitals are primarily non-bonding due to lower energy.

2. Assume greater overlap between N $p_x$ and O ($p_z + p_x$) compared to O ($p_y - p_y$) and N($s$) - making other $a_1$ orbitals non-bonding.

3. Assume $\sigma$ overlap > $\pi$ overlap.

\[ O - N = O \]

1 $\pi$ bond

2 $\sigma$ bonds

6 lone pair bonding

1 centered $N$

5 centered $O$

all consistent with MO diagram
If the molecule is linear (D_ooh) -

$\sigma_1 (p_y-p_y)$ is no longer $\sigma$ (in non-bonding)

changes to $\pi$ bonding

$b_2 (p_y+p_y)$ becomes non-bonding

$a_1, e'(p_z+nN)$ no longer involved in $\sigma$-bonding